

## 7.6 Review Questions

1. How is an isotope different than a neutral atom?

2. How is an ion different than a neutral atom?

3. Fill in the chart. You can round off the atomic numbers. The idea here is to try to fill in the line while knowing, at most, two things about the atom. If there isn't enough information to decide the last column, make it "vanilla". If you describe the atom as an isotope, tell if it is "fat" or "skinny". If you say the atom is an ion, list the overall charge (e.g. 1+ or 2-)

Symbol	Atomic Number	Atomic Mass	#P+	#No	#e-	Ion / Isotope/ or "vanilla"
Na						vanilla
	44	103				
			30	36		
B				5		
			63		62	
Au					79	
	9				10	
			20			2+ ion
					36	2- ion
Sc				24		